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## Kinetic effects of non-equilibrium plasma-assisted methane oxidation on diffusion flame extinction limits

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## ABSTRACT

The kinetic effects of low temperature non-equilibrium plasma assisted CH<sub>4</sub> oxidation on the extinction of partially premixed methane flames was studied at 60 Torr by blending 2% CH<sub>4</sub> by volume into the oxidizer stream of a counterflow system. The experiments showed that non-equilibrium plasma can dramatically accelerate the CH<sub>4</sub> oxidation at low temperature. The rapid CH<sub>4</sub> oxidation via plasma assisted combustion resulted in fast chemical heat release and extended the extinction limits significantly. Furthermore, experimental results showed that partial fuel mixing in the oxidizer stream led to a dramatic decrease of O concentration due to its rapid consumption by CH<sub>4</sub> oxidation at low temperature. The products of plasma assisted CH<sub>4</sub> oxidation were measured using the Two-photon Absorption Laser-Induced Fluorescence (TALIF) method (for atomic oxygen, O), Fourier Transform Infrared (FTIR) spectroscopy, and Gas Chromatography (GC). The product concentrations were used to validate the plasma assisted combustion kinetic model. The comparisons showed the kinetic model over-predicted the CO, H<sub>2</sub>O and H<sub>2</sub> concentrations and under-predicted CO<sub>2</sub> concentration. A path flux analysis showed that O generated by the plasma was the critical species for extinction enhancement. In addition, the results showed that O was produced mainly by direct electron impact dissociation reactions and the collisional dissociation reactions of electronically excited molecules with O<sub>2</sub>. Moreover, these reactions involving electron impact and excited species collisional dissociation of CH<sub>4</sub> contributed approximately a mole fraction of 0.1 of total radical production. The present experiments produced quantitative species and extinction data of low temperature plasma assisted combustion to constrain the uncertainties in plasma/flame kinetic models.

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## 1. Introduction

Recently, plasma assisted combustion has drawn attention for its potential to enhance combustion performance in gas turbines, scramjets, and internal combustion engines. Furthermore, extensive efforts have been made to understand the kinetic processes in plasma–flame interactions [1–17]. However, plasma assisted combustion involves strong coupling effects not only from the flow and heat interaction but also from the simultaneous production of active radicals, excited species, ions/electrons, and other intermediate species. Consequently, one of the major challenges is to decouple the complex plasma–flame interactions so that the kinetic contribution of individual species produced by plasma can be understood at the elementary reaction level [15,16]. As such, it is of great importance to (i) develop well-defined plasma-

combustion systems that can provide quantitative, fundamental experimental data to understand the effects of plasma generated species and (ii) develop and validate plasma-assisted combustion kinetic mechanisms.

Many experimental studies have been carried out to understand the role of plasma generated species on ignition, flame speed, and flame stabilization. For ignition studies, the reduction of ignition delay time by non-equilibrium nanosecond pulsed discharges was reported in Refs. [11,18]. The results showed that the ignition delay times could be reduced by about an order of magnitude with the action of plasma. An increase of atomic oxygen (O) concentration was observed with plasma and explained as a major factor for the decrease of ignition delay times. Recently, a major kinetic pathway of O production with plasma has been revealed through the quantitative measurements of O and excited N<sub>2</sub> concentrations [8,19], which has advanced the current understanding of plasma assisted combustion at the elementary reaction level.

The efforts on quantitative understanding of the enhancement mechanism through the flame/plasma interaction has been made

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by several research groups [12,15,16,20]. Ombrello and co-workers have demonstrated experimentally how to isolate singlet oxygen ( $O_2(a^1\Delta_g)$ ) and ozone ( $O_3$ ) from other plasma related species and shown that both (at concentrations of several thousand ppm) can enhance flame speeds by a few percent [15,16]. The experimental data further revealed that the current kinetic model for  $O_2(a^1\Delta_g)$  significantly over-predicted the enhancement for hydrocarbon flames, due to the lack of understanding of collisional quenching rates of  $O_2(a^1\Delta_g)$  with hydrocarbon species. Enhancement of flame stability by a non-equilibrium plasma discharge or an electric field in partially premixed systems also have been observed for lifted flames [13,14,21]. Additionally, different explanations for plasma assisted combustion have been given, based on the production of  $H_2$  and CO [13], the formation of active radicals such as OH and O [14], as well as the ionic wind [21]. Unfortunately, in spite of rigorous experimental investigations, the kinetic contribution from the plasma/flame interaction has not been fully characterized due to the complexity of either their interaction or the experiment itself.

Recently, a nanosecond pulsed discharge system has been integrated into the oxidizer side of a counterflow flame configuration, in order to obtain a simplified plasma/flame interaction [17]. The experimental results revealed that the extinction limit of diffusion flames can be significantly extended by applying the discharge on the oxidizer side to produce O. However, considering a practical application, where the flow condition is controlled by the intensive turbulent mixing, partial/full oxidation of the premixture with plasma would be one of the important factors that dominate the flame/plasma interactions. In this case, it is important to investigate how the chemical reactivity of the premixture can be induced or promoted through the plasma/flame interactions, particularly in the low temperature regime ( $<1000$  K). Moreover, due to a dearth of quantitative data, the current plasma combustion kinetic mechanism for ignition [8] remains untested for low temperature  $CH_4$  oxidation. Consequently, it is essential to provide quantitative measurements of species not only to validate of current kinetic models but also to advance the fundamental understanding of kinetic aspects of plasma/flame interactions.

Accordingly, the goal of the present study was to investigate the effect of plasma assisted  $CH_4$  oxidation at the low temperature regime on the extinction limits of diffusion flames with partially premixed oxidizer in the counterflow configuration. The species concentrations have been measured with TALIF for O and Fourier Transform Infrared (FTIR) spectroscopy and Gas Chromatography (GC) for stable intermediate species in order to characterize the performance of nanosecond pulsed discharge in a partially premixed oxidizer stream and to validate the current plasma-combustion kinetic model. This model has been validated further against the experimental measurements of extinction limits. Path flux analyses have also been conducted to identify and understand the important  $CH_4$  oxidation pathways in plasma assisted combustion.

## 2. Experimental methods and kinetic modeling

### 2.1. Partially premixed counterflow flame system with nano-second pulsed discharge

A schematic of the experimental system is shown in Fig. 1. The setup consisted of a pair of counterflow burners that were located in a low pressure chamber with a volume of 81 L. The inner diameters of the reactant and co-flow nozzles were 20 and 28 mm, respectively. The burner separation distance of the opposed nozzles were maintained to be equal to the nozzle diameter, 20 mm, to avoid the perturbation of the boundary condition from

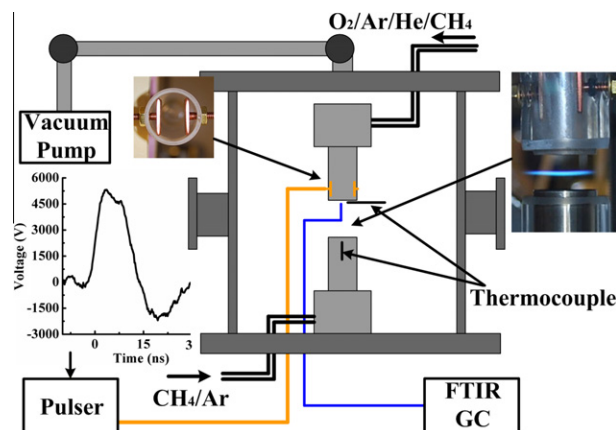


Fig. 1. Schematic of experimental setup.

thermal expansion. The co-flowing inert (Ar) curtain provided the isolation of the flame from the environment and burner cooling as well. A honeycomb plate was placed inside the reactant nozzles to ensure that the flow field was uniform. The counterflow nozzles were made of quartz, and two parallel bare metal electrodes (a copper alloy that is corrosion resistant to avoid the catalytic effect of copper) were located inside the tubes of the upper nozzle. The electrodes were  $15\text{ mm} \times 22\text{ mm} \times 1\text{ mm}$  (thick) and were separated by 10 mm (inset in Fig. 1). The electrode edges were smoothed to avoid the local concentration of electric field and dielectric breakdown. The front ends of the electrodes were 10 mm away from the exit to avoid electrical interaction with the thermocouples. The experimental pressure was 60 Torr, the diffusion time scale (approximately 1 ms, estimated by  $l^2/D$ , where  $l$  was the electrode thickness, and  $D$  the diffusivity ( $\sim 10^{-3}\text{ m}^2/\text{s}$ ) of  $O_2$  at 300 K and 60 Torr) was much shorter than the flow residence time (approximately 10 ms). Thus, the diffusion was sufficiently fast, and the effect of the electrodes on the flow field inside the tube was negligible. To ensure that the stagnation plane was formed in the middle of the two burner exits, the momenta of the two reactant streams were matched. During the experiments, He and Ar were used as dilution gases for the partially premixed oxidizer stream (He can improve the uniformity of the discharge [22], but it was difficult to maintain the flame if He was used as the only dilution gas due to the large thermal diffusivity at low pressure. Thus in our experiments, a He and Ar mixture was used as the diluent). The composition of the oxidizer stream was fixed at  $O_2/Ar/He/CH_4$  (0.26:0.32:0.4:0.02) to observe the effect of  $CH_4$  doping in the oxidizer. The fuel stream was  $CH_4$  diluted by Ar (fuel mole fraction varied from 0.2 to 0.4). A nanosecond pulse generator (FID FPG 30-50MC4) was used to generate the non-equilibrium plasma to activate the partially premixed oxidizer stream. The pulse generator is capable of producing repetitive 32 kV pulses with pulse duration of 12 ns, full width at half maximum (FWHM). The frequency of the pulse generator is adjustable from 1 to 50 kHz. A typical voltage waveform measured by a LeCroy high voltage probe (PPE20KV), is embedded in Fig. 1. The current through the electrodes was measured with a Pearson Coil (Model 6585). The pulse energy supplied to the discharge was estimated from the time integration of the voltage and current profiles and found to be approximately 1.27 mJ/pulse. The characteristics of voltage–current were independent of pulse repetition frequency. Therefore, average powers were calculated to be 5.1, 12.7, 25.4, 38.1 and 50.8 W at the respective frequencies of 4, 10, 20, 30 and 40 kHz. The amplitude of the voltage was fixed at 5.3 kV for all tested experimental conditions, as shown in Fig. 1. The direct image of the nanosecond repetitive discharge is shown in Fig. 2a at

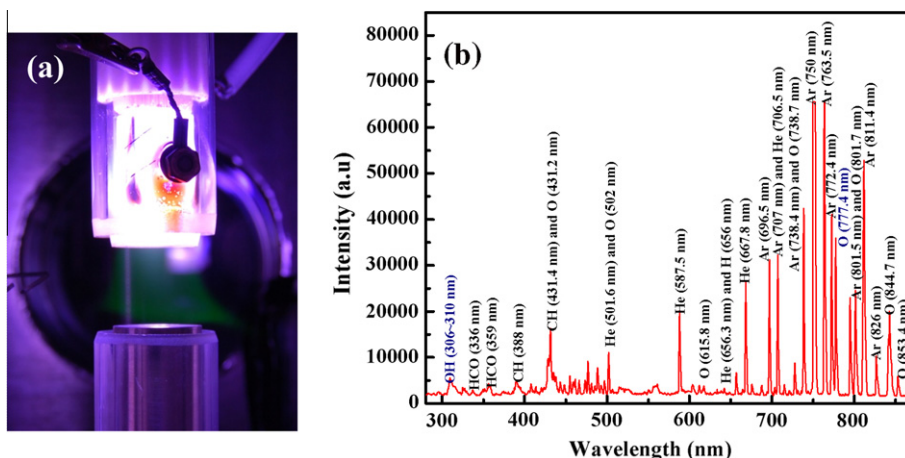


Fig. 2. Discharge at Ar/He/O<sub>2</sub>/CH<sub>4</sub> (0.32/0.4/0.26/0.02 by volume),  $d = 20$  mm,  $f = 30$  kHz,  $P = 60$  Torr, (a) direct image, (b) optical emission spectrum.

a pulse repetition frequency of 30 kHz. The optical emission spectrum of the discharge is shown in Fig. 2b. Figure 2b shows strong emissions from excited O, OH, HCO, CH, He, and Ar, which indicated the radical formation from CH<sub>4</sub> oxidation during the discharge. It was also observed that the strongest emissions were from excited Ar and O (750 nm and 777.4 nm, respectively) and the intensity of all emission lines increased with the increase of pulse repetition frequency. Thus, the plasma assisted CH<sub>4</sub> oxidation was presumed to be affected by the pulse repetition frequency, and the excited Ar and O may play an important role in the CH<sub>4</sub> oxidation. The intensity of excited He emission was not strong, however; this is because the excitation energy of He is relatively high (19.8 eV) compared to that of Ar (11.5 eV). Detailed discussions of the kinetic process will be given in the next section.

The extinction strain rates were measured by fixing the fuel mole fraction and increasing the flow velocity gradually until extinction occurred. At the same time, temperatures at the burner exits were measured simultaneously using thermocouples. One thermocouple (380  $\mu$ m bead diameter) was fixed at the exit of the lower fuel nozzle, and another movable thermocouple was used to monitor the temperature at the exit of the upper nozzle. The movable thermocouple was coated with MgO and covered with a grounded Nickel–Chrome sheath (with 1 mm outer diameter) to remove the electromagnetic interference from the discharge.

The major products (CH<sub>4</sub>, CO, CO<sub>2</sub>, CH<sub>2</sub>O, and H<sub>2</sub>O) from the plasma assisted CH<sub>4</sub> oxidation in the oxidizer stream were measured by using an FTIR spectrometer (Nicolet Magna-IR 550). A quartz probe with an inlet diameter 1 mm was placed axially at the oxidizer side nozzle exit and was attached to a heated line (to avoid H<sub>2</sub>O condensation) of the FTIR spectrometer. The pressure and temperature in the FTIR system were held constant at 31 Torr and 393 K, respectively. For H<sub>2</sub> measurement, a micro GC system (INFICON 3000) was used. The calibrations were performed by flowing sample gas with known species concentrations through the FTIR or GC at conditions identical to those of the experiments. The relative mole fraction uncertainties of sampling measurements were found to be less than  $\pm 1\%$  for CH<sub>4</sub>, CO, CO<sub>2</sub> and  $\pm 5\%$  for H<sub>2</sub>O and H<sub>2</sub>. The uncertainty of the CH<sub>2</sub>O measurement was  $\pm 80$  ppm. The sampling was also performed by using a quartz probe with 6 mm diameter. The results were identical with that of 1 mm diameter probe which indicated the uniformity of the properties of the counterflow system in the center region and the validity of one dimensional approximation.

## 2.2. Two-photon Absorption Laser Induced Fluorescence (TALIF) system

The TALIF method, calibrated with Xenon, was used to measure the absolute O concentration at the oxidizer-side nozzle exit. The TALIF system used in this study was described in detail in our previous study [17], and thus only a summary of this method is presented here. Ground state O is excited by absorbing two photons at a wavelength of 225.7 nm. The transition between the excited  $3p^3P$  state and the  $3s^3S$  state will release a single photon at 844.6 nm. Xenon can be excited from  $5p^6\ ^1S_0$  to  $6p^1\ [3/2]2$  with the two photons at 224.31 nm; de-excitation to  $6s^1\ [1/2]1$  corresponds to fluorescence at 834.91 nm. The number density of O ( $N_O$ ) was calculated using the following equation in terms of the known number density of Xenon ( $N_{Xe}$ ) [8],

$$N_O = \frac{S_O}{S_{Xe}} g_{ND} \frac{a_{21}(Xe)}{a_{21}(O)} \left( \frac{\sigma^{(2)}(Xe)}{\sigma^{(2)}(O)} \right) \left( \frac{v_O}{v_{Xe}} \right)^2 \times \frac{1}{F_O(T)} N_{Xe}$$

where  $S_O$  and  $S_{Xe}$  are the observed fluorescence signals for O and Xenon, respectively,  $a_{21} = \frac{A_{21}}{A_{21} + Q}$  the fluorescence quantum yields ( $A_{21}$  and  $Q$  are spontaneous emission and quenching rates, respectively),  $\sigma^{(2)}$  the two photon absorption cross sections of Xe and O,  $F_O(T)$  the atomic oxygen Boltzmann factor for the lower level of the two photon absorption,  $v_i$  the photon energies, and  $g_{ND}$  the neutral density filter factor ( $1.82 \times 10^{-3}$  from manufacturer). These values can be found or easily calculated based on the data in Refs. [8,23].

An Nd:YAG laser was used to generate 532 nm to pump a tunable dye laser operating at  $\sim 573$  nm. This 573 nm beam was frequency doubled and mixed with the 1064 nm beam of the Nd:YAG laser to get a  $\sim 226$  nm beam of  $\sim 10$   $\mu$ J/pulse required for the TALIF diagnostics. The TALIF signal was observed using an 850 nm bandpass filter of 40 nm FWHM and a Hamamatsu photomultiplier (R636-10). A separate vacuum cell was used to obtain the Xenon atom two-photon fluorescence spectrum. Corrections for the transmission of the vacuum cell and bandpass filter were made to find the O density by the method described in Refs. [8,23]. The overall relative mole fraction uncertainty of the TALIF measurement was approximately  $\pm 40\%$ .

## 2.3. Computational approach with plasma-combustion kinetic model

The kinetic model used in this work was an integration of the air plasma model [8] together with USC Mech II [24]. In addition,



in order to consider species dissociation by excited Ar and He and dissociation of  $\text{CH}_4$  by electron impact, additional elementary reactions involving the dilution gases (He/Ar) and  $\text{CH}_4$  were added (Table 1). The plasma combustion model incorporates a set of species conservation equations for number densities of neutral, charged, and excited species produced in the plasma,  $\text{O}$ ,  $\text{O}_2$ ,  $\text{O}_3$ ,  $\text{e}$ ,  $\text{O}_2^+$ ,  $\text{O}^+$ ,  $\text{O}^-$ ,  $\text{O}_2^-$ ,  $\text{O}_2(\text{a}^1\Delta_g)$ ,  $\text{O}_2(\text{b}^1\Sigma)$ ,  $\text{O}(\text{D})$ ,  $\text{Ar}(+)$ ,  $\text{Ar}(3\text{p}^5(^2\text{P}_{3/2})4\text{s})$ ,  $\text{Ar}(3\text{p}^5(^2\text{P}_{1/2})4\text{s})$ ,  $\text{Ar}(3\text{p}^5(^2\text{P}_{3/2})4\text{p})$  ( $\text{Ar}^*$  for simplicity, with respective excitation energies of 11.5, 11.7, and 12.9 eV),  $\text{He}(+)$ ,  $\text{He}(2\text{s}^3\text{S})$ ,  $\text{He}(2\text{s}^1\text{S})$ ,  $\text{He}(2\text{p}^3\text{P})$ ,  $\text{He}(2\text{p}^1\text{P})$ ,  $\text{He}(3\text{s}^1\text{S})$ , and  $\text{He}(4\text{p}^1\text{P})$  ( $\text{He}^*$  for simplicity, with respective excitation energies of 19.8, 20.6, 21, 21.2, 22.9, and 23.7 eV), as well as the energy conservation

**Table 1**  
List of Ar/He/ $\text{CH}_4$  related reactions involved in Ar/He/ $\text{O}_2$ / $\text{CH}_4$  mixture discharge.

Reaction	Rate ( $\text{cm}^3 \text{s}^{-1}$ )	References
$\text{e} + \text{CH}_4 = \text{e} + \text{CH}_3 + \text{H}$	$\sigma^a$	[27]
$\text{e} + \text{CH}_4 = \text{e} + \text{CH}_4(+)$	$\sigma$	[28]
$\text{e} + \text{CH}_4(+) = \text{CH}_2 + 2\text{H}$	$1.7 \times 10^{-7} (300/T)^{0.5}$	[28]
$\text{e} + \text{CH}_4(+) = \text{CH}_3 + \text{H}$	$1.7 \times 10^{-7} (300/T)^{0.5}$	[28]
$\text{CH}_4(+) + \text{O}_2 = \text{CH}_4 + \text{O}_2(+)$	$5 \times 10^{-10}$	[28]
$\text{Ar} + \text{e} = \text{Ar}^* + \text{e}$	$\sigma$	[28]
$\text{Ar} + \text{e} = \text{Ar}(+) + 2\text{e}$	$\sigma$	[28]
$\text{Ar}^* + \text{O}_2 = \text{Ar} + 2\text{O}$	$2 \times 10^{-10}$	[28]
$\text{Ar}(+) + \text{O}_2 = \text{Ar} + \text{O}_2(+)$	$1 \times 10^{-10}$	[28]
$\text{Ar}^* + \text{CH}_4 = \text{Ar} + \text{CH}_2 + 2\text{H}$	$3.3 \times 10^{-10}$	[28]
$\text{Ar}^* + \text{CH}_4 = \text{Ar} + \text{CH} + \text{H}_2 + \text{H}$	$5.8 \times 10^{-10}$	[28]
$\text{Ar}^* + \text{CH}_4 = \text{Ar} + \text{CH}_3 + \text{H}$	$5.8 \times 10^{-10}$	[28]
$\text{Ar}^* + \text{CH}_4 = \text{Ar} + \text{CH}_2 + \text{H}_2$	$5.8 \times 10^{-10}$	[28]
$\text{Ar}(+) + \text{CH}_4 = \text{Ar} + \text{CH}_3(+) + \text{H}$	$6.5 \times 10^{-10}$	[28]
$\text{Ar}(+) + \text{CH}_4 = \text{Ar} + \text{CH}_2(+) + \text{H}_2$	$1.4 \times 10^{-10}$	[28]
$\text{He} + \text{e} = \text{He}^* + \text{e}$	$\sigma$	[29]
$\text{He} + \text{e} = \text{He}(+) + 2\text{e}$	$\sigma$	[29]
$\text{He}^* + \text{O}_2 = \text{He} + \text{O}_2(+) + \text{e}$	$1.5 \times 10^{-11} T^{0.5}$	[29]
$\text{He}^* + \text{O} = \text{He} + \text{O}(+) + \text{e}$	$1.5 \times 10^{-11} T^{0.5}$	[29]
$\text{He}(+) + \text{O}_2 = \text{He} + \text{O}(+) + \text{O}$	$0.6 \times 10^{-11} T^{0.5}$	[29]
$\text{He}(+) + \text{O}_3 = \text{He} + \text{O}_2 + \text{O}(+)$	$0.6 \times 10^{-11} T^{0.5}$	[29]
$\text{He}(+) + \text{O}_2(\text{a}^1\Delta_g) = \text{He} + \text{O}(+) + \text{O}$	$0.6 \times 10^{-11} T^{0.5}$	[29]
$\text{He}(+) + \text{O}(\text{D}) = \text{He} + \text{O}(+)$	$2.9 \times 10^{-12} T^{0.5}$	[29]
$\text{He} + 2\text{O} = \text{He}^* + \text{O}_2$	$1 \times 10^{-33}$	[29]
$\text{He} + \text{O}(\text{D}) = \text{He} + \text{O}$	$1 \times 10^{-10}$	[30]
$\text{He}^* + \text{CH}_4 = \text{He} + \text{CH} + \text{H}_2 + \text{H}$	$5.6 \times 10^{-13}$	[31]

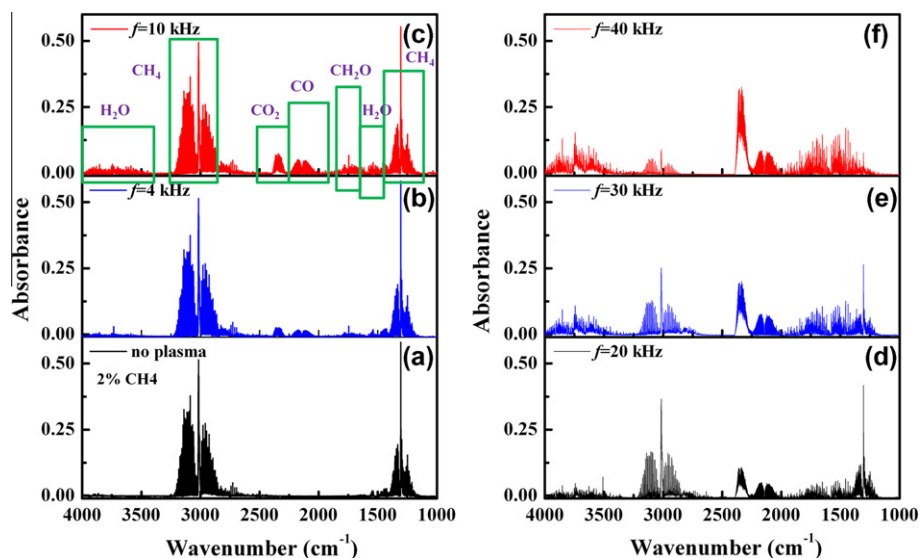
<sup>a</sup> The rate is calculated by the Boltzmann solver using experimentally measured cross-sections,  $\sigma$ .

equation for prediction of the temperature of the mixture. This set of equations are coupled with the steady state, two-term expansion Boltzmann equation for the electron energy distribution function (EEDF) of the plasma electrons using experimentally measured cross sections of electron impact electronic excitation, dissociation, ionization, and dissociative attachment processes. The Boltzmann equation calculates the rate coefficients of the electron impact elementary reactions by averaging the cross sections over the EEDF. The kinetic model also incorporates chemical reactions of excited electronic species, electron–ion recombination and ion–ion neutralization processes, ion–molecule reactions, and electron attachment and detachment processes. Note that the present model does not solve the Poisson equation for the electric field and therefore does not take into account charge separation and sheath formation near the electrode. So in the present work, the reduced electrical field ( $E/N$  ratio) in the plasma was considered to be an adjustable parameter, varied until the  $\text{CH}_4$  concentration (at the nozzle exit, 1 cm away from the end of the discharge region) matched with the experimental results at the burner exit.  $\text{CH}_4$  concentration was used here to determine the  $E/N$  ratio because  $\text{CH}_4$  concentration measurement had the smallest uncertainty (1% vs. 40% for TALIF of O). Therefore, only the concentrations of CO,  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ ,  $\text{H}_2$ ,  $\text{CH}_2\text{O}$ , and O from the simulations were compared with the experiments. The model was also employed to calculate the radical/atom (OH and H) concentrations by matching the O concentrations from the simulations with the corresponding experimental values. The radical concentrations were further used as the boundary conditions for extinction limit calculations in Section 3.2. More details about this model can be found at Refs. [25,26].

### 3. Results and discussion

#### 3.1. Effects of discharge repetition rate on species concentrations from plasma assisted $\text{CH}_4$ oxidation

FTIR and GC instrumentation were used to quantify the major stable products of the low temperature plasma assisted  $\text{CH}_4$  oxidation by sampling at the center of the nozzle exit in the oxidizer stream using quartz probe. Typical FTIR absorption spectra are shown in Fig. 3a–f. Figure 3a shows the FTIR absorption spectra



**Fig. 3.** FTIR spectrum with different pulse repetition frequencies at Ar/He/ $\text{O}_2$ / $\text{CH}_4$  (0.32/0.4/0.26/0.02 by volume),  $P = 60$  Torr; (a) no plasma, (b)  $f = 4$  kHz, (c)  $f = 10$  kHz, (d)  $f = 20$  kHz, (e)  $f = 30$  kHz, and (f)  $f = 40$  kHz.

without the discharge. The  $\text{CH}_4$  absorbance spectrum is seen clearly and shows no effects of any  $\text{CO}_2$  and  $\text{H}_2\text{O}$  entrainment from the environment. With the presence of the plasma discharge, the absorbance spectra of  $\text{CO}$ ,  $\text{CO}_2$ ,  $\text{H}_2\text{O}$  and  $\text{CH}_2\text{O}$  appeared due to the plasma assisted  $\text{CH}_4$  oxidation. Moreover, as shown in Fig. 3b–f, with the increase of discharge pulse repetition frequency (or equivalently, discharge power), the concentrations of  $\text{CO}$ ,  $\text{CO}_2$  and  $\text{H}_2\text{O}$  increased significantly, and at the same time, the concentration of  $\text{CH}_4$  decreased. The dependence of the measured concentrations of  $\text{CO}$ ,  $\text{CO}_2$ ,  $\text{H}_2$ ,  $\text{H}_2\text{O}$ , and  $\text{CH}_2\text{O}$  on the plasma pulse repetition rate is shown in Figs. 4–6. In order to validate the plasma flame model, the predicted species concentrations are also plotted in Figs. 5 and 6. In all the FTIR measurements, it was confirmed that the carbon closure with the measured species was 95%, demonstrating that the measurements captured the major carbon containing species in the plasma assisted  $\text{CH}_4$  oxidation. It was also observed that carbon black deposited on the electrodes and sampling quartz probe during the experiments and that the minor carbon loss was caused by the formation of carbon black in the plasma discharge. As shown in Figs. 4 and 5, with the increase of pulse repetition frequency, the  $\text{CH}_4$  concentration decreased significantly, and nearly a mole fraction of 0.9 of  $\text{CH}_4$  was oxidized at pulse repetition frequency  $f = 40$  kHz. At the same time, the concentrations of  $\text{CO}$ ,  $\text{CO}_2$  and  $\text{H}_2\text{O}$  increased monotonically with  $f$ . With the increase of  $\text{CH}_4$  oxidation ratio, the oxidizer temperature at the nozzle exit also increased significantly ( $423 \pm 3$  K,  $613 \pm 5$  K,  $743 \pm 5$  K,  $843 \pm 5$  K and  $933 \pm 8$  K at  $f = 4$  kHz, 10 kHz, 20 kHz, 30 kHz, and 40 kHz, respectively). Note that the concentration of  $\text{H}_2$  after low

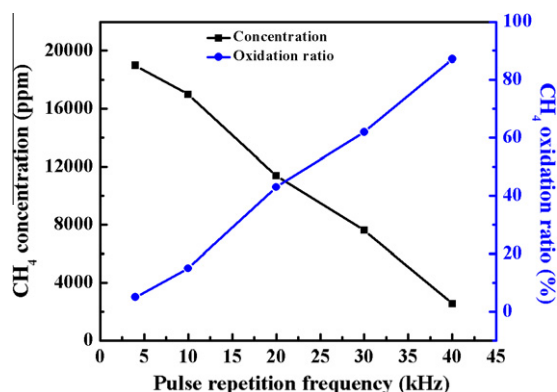


Fig. 4. Dependence of  $\text{CH}_4$  concentrations and oxidization ratios on pulse repetition frequency.

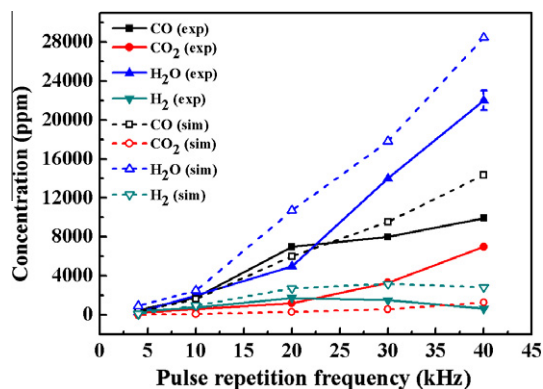


Fig. 5. Dependence of species concentrations at the burner exit on pulse repetition frequency (experiments: solid symbol and solid line; simulations: open symbol and dashed line).

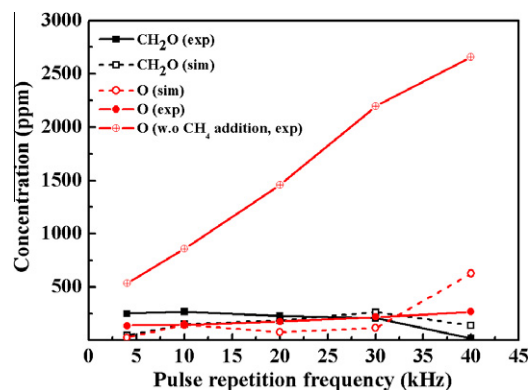
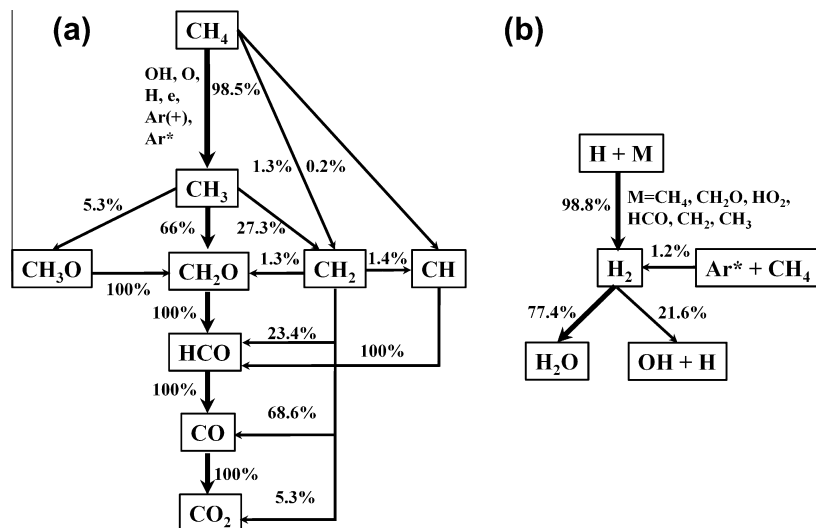
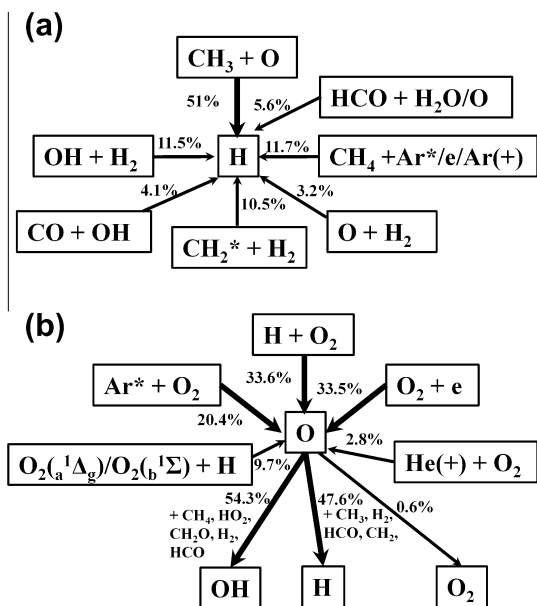


Fig. 6. Dependence of species concentrations at the burner exit on pulse repetition frequency (experiments: solid symbol and solid line; simulations: open symbol and dashed line).

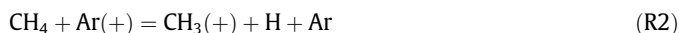
temperature plasma assisted  $\text{CH}_4$  oxidation remained very low (less than 2000 ppm). The peak concentration of  $\text{H}_2$  occurred at a pulse repetition frequency  $f = 20$  kHz. The low value of peak  $\text{H}_2$  concentration compared to  $\text{H}_2\text{O}$  concentration suggests that the plasma assisted combustion was not dominated by fuel reforming into  $\text{H}_2$  and  $\text{CO}$ , as will be discussed later. A low peak concentration for  $\text{CH}_2\text{O}$  at  $f = 10$  kHz was also observed during the experiments, as shown in Fig. 6. It is shown in both Figs. 5 and 6 that the model captured the correct trend of the relationship between the species concentrations and pulse repetition frequency, but the kinetic model over-predicted  $\text{CO}$  (up to 44%),  $\text{CH}_2\text{O}$  (up to a factor of 6) and  $\text{H}_2$  (up to a factor of 3), and under-predicted  $\text{CO}_2$  (up to 89%) by volume, indicating that the model underestimated the overall reactivity in the oxidizer stream.

In order to identify the important  $\text{CH}_4$  oxidation pathways and understand the kinetic processes in the plasma assisted  $\text{CH}_4$  oxidation mechanism, path flux analysis was performed for a pulse repetition frequency  $f = 40$  kHz, and the results are shown in Fig. 7. Methane was dissociated to  $\text{CH}_3$  mainly by H abstraction with OH, O and H produced via the plasma discharge. Methane also decomposed to  $\text{CH}_3$ ,  $\text{CH}_2$  and  $\text{CH}$  via collisions with an electron,  $\text{Ar}^+$  and  $\text{Ar}^*$  (reactions (1)–(6)). The formed  $\text{CH}_3$ ,  $\text{CH}_2$  and  $\text{CH}$  were oxidized further to  $\text{CH}_2\text{O}$ ,  $\text{HCO}$ ,  $\text{CO}$  and finally  $\text{CO}_2$  as shown in Fig. 7a. Among those reaction paths, a total of a mole fraction of 0.12 of  $\text{CH}_4$  was decomposed by collisions with electrons,  $\text{Ar}^+$  and  $\text{Ar}^*$ . Molecular hydrogen is another important product of plasma assisted  $\text{CH}_4$  oxidation because it can affect greatly the ignition and extinction processes via its fast oxidation chemistry. The reaction path of  $\text{H}_2$  is shown in Fig. 7b. The dominant formation path of  $\text{H}_2$  was through H abstraction reactions with  $\text{CH}_4$ ,  $\text{CH}_2\text{O}$ ,  $\text{HO}_2$ ,  $\text{HCO}$ ,  $\text{CH}_2$  and  $\text{CH}_3$ . The formation of  $\text{H}_2$  via  $\text{H} + \text{HO}_2 = \text{H}_2 + \text{O}_2$  clearly indicates the nature of low temperature chemistry. This reaction plays an important role at low temperature and high pressure when  $\text{HO}_2$  concentration is high [32]. Another formation path of  $\text{H}_2$  was through  $\text{CH}_4$  decomposition by  $\text{Ar}^*$ , but only 1.2% of  $\text{H}_2$  was formed through this reaction path. Moreover, the concentration of  $\text{H}_2$  increased at low pulse repetition frequency, due to the increased H abstraction reaction by H radicals; at high frequency,  $\text{H}_2$  concentration then decreased via increased reactivity at higher temperature. Therefore, the concentration of H is important to determining the concentration of  $\text{H}_2$ . The reaction paths of H and O are also shown in Fig. 8a and b. Only 11.7% of H was formed via the decomposition of  $\text{CH}_4$  by electron impact and excited species collisions through reactions (R1), (R2), (R3), (R4), (R5). The major contribution to H formation was from reaction  $\text{CH}_3 + \text{O} = \text{CH}_2\text{O} + \text{H}$ . Thus, O concentration is crucial to determining H concentration.

Fig. 7. Reaction paths flux analysis for (a) CH<sub>4</sub> and (b) H<sub>2</sub> at  $f = 40$  kHz.Fig. 8. Reaction paths flux analysis for (a) H and (b) O at  $f = 40$  kHz.

tion are formed from the reaction of O with CH<sub>3</sub>, which is itself produced by the direct plasma discharge in the partially premixed CH<sub>4</sub>/O<sub>2</sub> stream. Therefore, it can be concluded that O production by plasma is the major source to initiate the low temperature plasma assisted CH<sub>4</sub> oxidation. Moreover, the partial CH<sub>4</sub>/O<sub>2</sub> premixing increases H production and accelerates the chain branching process. Once O is generated by the discharge, it is consumed rapidly by CH<sub>4</sub> and its intermediate oxidized and dissociated products (CH<sub>3</sub>, CH<sub>2</sub>, CH<sub>2</sub>O, HCO, HO<sub>2</sub> and H<sub>2</sub>). It is shown in Fig. 8b that the major consumption path of O is to generate OH and H, which are important to the CH<sub>4</sub> oxidation process. OH and H then further react with fuel and its fragments, which become the dominant reaction paths of the fuel decomposition. Only a small amount (0.6%) of O recombines to O<sub>2</sub> (Surface recombination reaction of O was not considered in this study. This was accounted for by measuring the O concentration at the center of the nozzle and comparing with the simulations). Note that the reaction path from O to O<sub>3</sub> is negligible in this study due to the high temperature conditions. Furthermore, the concentrations of reactive oxidant ions (O<sub>2</sub><sup>+</sup>, O<sub>2</sub><sup>-</sup>) are several orders of magnitude lower than O, and thus their effects are also negligible. This indicates that when the discharge is applied to the premixture of fuel and oxidizer, O can be consumed more efficiently by the fuel and its intermediate oxidation and dissociation products rather than through recombination. Based on the above analysis, it is clear that O production within a plasma discharge in a partially premixed fuel and oxidizer is critical for H<sub>2</sub> and H formation and subsequent CH<sub>4</sub> oxidation. Therefore, in order to predict other species, O formation through the collision-induced O<sub>2</sub> dissociation by electrons and excited species needs to be accurately predicted. In addition, to improve the CH<sub>4</sub> oxidation efficiency (the consumption efficiency of O), the plasma discharge needs to be implemented in a partially premixed lean fuel/oxidizer mixture.

In order to determine the concentration of O and compare with the kinetic model prediction, the Xenon calibrated, Two Photon Absorption Laser Induced Fluorescence (TALIF) method was used. The relationship between O concentration and pulse repetition frequency (for a flow velocity of 200 cm/s) with and without fuel premixing is plotted in Fig. 6. For plasma discharge in a partial premixture of fuel and oxidizer, it is shown that with the increase of pulse repetition frequency, the O concentration increased, but the increase was very low, less than 300 ppm even at  $f = 40$  kHz (note that the TALIF detection limit can be as low as several ppm



The reaction path of O is shown in Fig. 8b. Approximately 54% of O was formed through collisions between O<sub>2</sub> and electrons/Ar\*. Among all the three major O production paths, 33.5% was from direct electron impact dissociation of O<sub>2</sub>, 20.4% from O<sub>2</sub> collisions with Ar\*, and 33.6% from the H + O<sub>2</sub> = OH + O branching reaction. Note from Fig. 8a that the initial H radicals for the branching reac-



[8]). The low rate of increase of O production versus  $f$  suggests that the cause is rapid CH<sub>4</sub> oxidation at high plasma repetition frequency by the induced chain reactions (shown in Fig. 8b). To demonstrate this fact, the O concentration without fuel premixing was also measured at the same condition. Figure 6 shows that the concentration of O increased sharply from 530 ppm to 2700 ppm when the pulse repetition frequency increased from 4 kHz to 40 kHz. This result confirmed that the increased CH<sub>4</sub> oxidation at high plasma frequency was due to the increased production of O. The simulation results of O and CH<sub>2</sub>O concentrations are also shown in Fig. 6, together with the experimental measurements of CH<sub>2</sub>O concentration. The calculated concentrations of O and CH<sub>2</sub>O intersected with the experimental results, but the deviations became large at  $f = 40$  kHz.

As analyzed, the model under-predicted the concentration of CO<sub>2</sub>, and over-predicted CO and H<sub>2</sub>, which indicated an overall lower reactivity of the mixture. For O and CH<sub>2</sub>O prediction, the kinetic model significantly under-predicted their concentrations at low pulse repetition frequency and over-predicted the concentrations at high pulse repetition frequency. This may be due to the missing reaction paths involving peroxide species at low temperature and those involving interactions of excited species with hydrocarbon species and its products. For example, the HO<sub>2</sub> + X (X = H, OH) reaction rates have large uncertainty at low temperature, and the OH + O + M = HO<sub>2</sub> + M reaction, which is not included in the current combustion mechanism, may become important at low temperature [32]. In addition, a fraction of O atoms was excited to O(<sup>1</sup>D), which is much more reactive, especially at low temperature conditions. But, unfortunately, reactions involving O(<sup>1</sup>D) and hydrocarbon species are not well characterized and are absent in the current kinetic model (except O(<sup>1</sup>D) + H<sub>2</sub> = H + OH [33], which is 10<sup>7</sup> faster than O + H<sub>2</sub> = H + OH at 300 K). The formation of CH<sub>2</sub>O was primarily from the reaction CH<sub>3</sub> + O = CH<sub>2</sub>O + H. If the concentration of O was over-predicted (or under-predicted), the concentration of CH<sub>2</sub>O would also be over-predicted (or under-predicted). Likewise, H<sub>2</sub>O can react with O(<sup>1</sup>D) via the H<sub>2</sub>O + O(<sup>1</sup>D) = 2OH reaction and can also be dissociated by electrons and Ar\* to generate OH and H. The generation of OH can accelerate the oxidation of CO.

The kinetic model also showed the formation of C<sub>2</sub>H<sub>6</sub> (~200 ppm), but the C<sub>2</sub>H<sub>6</sub> absorption spectrum was not observed in the FTIR experiments. Previous research in a lean CH<sub>4</sub>/air discharge [34] also showed no C<sub>2</sub>H<sub>6</sub> formation, and thus the reaction CH<sub>3</sub> + CH<sub>3</sub> = C<sub>2</sub>H<sub>6</sub> was over-estimated or the consumption path of C<sub>2</sub>H<sub>6</sub> was underestimated at low temperatures. Furthermore, it is known that the current kinetic mechanism has not been validated below 700 K. More quantitative experimental data at low temperature conditions are thus required to improve the predictability of the kinetic mechanism.

### 3.2. Extinction strain rates measurement and computation

The extinction strain rates of the CH<sub>4</sub>/O<sub>2</sub> diffusion flames were measured without and with plasma (at  $f = 4, 10, 20, 30$ , and 40 kHz). A definition of global strain rate [35]

$$a_0 = \frac{2U_o}{L} \left( 1 + \frac{U_f \sqrt{\rho_f}}{U_o \sqrt{\rho_o}} \right) \quad (1)$$

was adopted and compared with numerical simulations. Here  $U$  is the speed of the flow,  $\rho$  the density of the flow, and  $L$  the gap distance between the two nozzles. The subscripts o and f refer to the oxidizer and fuel side, respectively. The use of global strain rate was found to be appropriate for mechanism validation [36], and the global strain rates were used in this study for both experiments and simulations.

In this experiment, the composition of the partially premixed oxidizer stream was fixed at a volume ratio of 0.02/0.26/0.32/0.4 for CH<sub>4</sub>/O<sub>2</sub>/Ar/He. The fuel stream was composed of CH<sub>4</sub> and Ar, and the CH<sub>4</sub> mole fractions varied from 0.2 to 0.4. During the measurements, the fuel mole fraction was fixed, and the flow velocity of both sides increased gradually. With the increase of flow velocity, the strain rate increased, and the flame had less residence time to complete the CH<sub>4</sub> oxidation and thus extinguished when the strain rate was above a critical value. This critical strain rate at flame extinction was recorded as the global extinction strain rate. Figure 9 shows the relationship of extinction strain rates and fuel mole fractions without plasma and with plasma at pulse repetition frequencies of  $f = 4$  and 10 kHz (with oxidizer temperature  $T_o = 423 \pm 4$  K and  $613 \pm 5$  K, the oxidizer side flow rates were among 1.92–2.18 SLPM and 2.22–2.41 SLPM at fuel mole fractions from 0.34 to 0.4, respectively). Extinction strain rates without CH<sub>4</sub> addition (O<sub>2</sub>/Ar/He ratio of 0.28/0.32/0.4) were also presented in Fig. 9 as reference. The experiments were compared further with numerical simulations. The simulations were computed using OPP-DIFF of the CHEMKIN package [37] with a modified arc-length continuation method [38] for plug flow. All simulations were performed by setting the boundary conditions using the measured temperatures and species concentrations (using FTIR, GC and TALIF instruments). The concentrations of H and OH were estimated from the simulation by matching the computational and experimental values of O concentrations, as discussed in Section 3.1. It was noticed that the extinction strain rate was sensitive to the CH<sub>4</sub> mole fraction in the oxidizer stream because this determines the heat release rate in the reaction zone. Thus, the failure to update the species concentrations information as the boundary condition may cause approximately 15% larger calculated extinction strain rates in the simulations. The results in Fig. 9 show that with CH<sub>4</sub> addition, the extinction strain rates were extended significantly and the oxidizer temperature after the plasma activation was higher (613 K and 548 K with and without CH<sub>4</sub> addition at  $f = 10$  kHz, respectively). It is also shown that with the increase of the pulse repetition frequency,  $f$ , there was a significant increase of extinction limit enhancement and the measured extinction strain rates agree well with the computed values. The present results indicate that the plasma generated CH<sub>4</sub> oxidation plays an important role for the extension of flame extinction limit. This is because when the flame approaches extinction, there was not enough residence time for the CH<sub>4</sub> to complete the oxidation reactions to release the chemical heat. However, plasma can dramatically accelerate the CH<sub>4</sub> oxidation to release the chemical energy at low temperature. The prompt chemical heat release extended the extinction limit.

It was noticed that during the plasma assisted CH<sub>4</sub> oxidation process, part of the CH<sub>4</sub> was reformed into H<sub>2</sub> and CO. Molecular

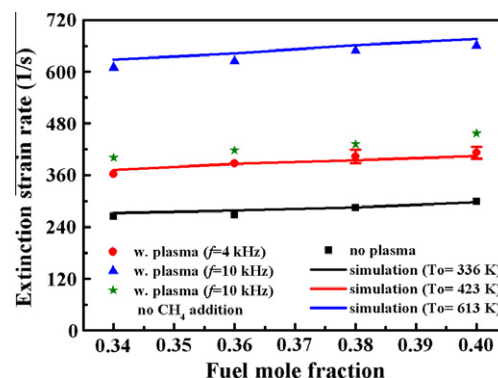


Fig. 9. Dependence of extinction strain rate on fuel mole fraction at the fuel side under different experimental conditions.

hydrogen can extend the extinction limit through its fast chemistry, and previous research [13,20,39] showed that plasma induced fuel reforming, producing a  $H_2/CO$  mixture, can enhance the flame stability. Therefore, it is important to examine how the fuel reforming affects the extinction. The present experiment showed that the  $H_2$  formation by non-equilibrium plasma discharge was low. As such, the plasma reforming effect only played a minor role in the extension of the extinction limit. Of course, the  $H_2$  yield depends on the discharge characteristics and fuel concentration in the mixture. It is necessary to understand whether fuel reforming or  $CH_4$  oxidation is more effective in increasing flame stabilization.

Because plasma can cause both  $CH_4$  oxidation and reformation, simulations of the individual effects of plasma assisted  $CH_4$  oxidation and  $CH_4$  reforming were performed and compared in Fig. 10. Here, part of the fuel ( $CH_4$ ) in the oxidizer stream was assumed to be ideally oxidized to  $H_2O$  and  $CO_2$  for Case 1 and reformed to  $H_2$  and  $CO$  in Case 2 under a constant enthalpy constraint. It is shown that extinction was enhanced for both cases, but the enhancement from the plasma assisted  $CH_4$  oxidation (Case 1) is much larger than that of  $CH_4$  reforming. This is because when the flame was approaching extinction, the residence time was too short to complete the reactions. That is,  $CH_4$  and its reformation products in the oxidizer stream cannot react completely to release the chemical enthalpy in the reaction zone in Case 2, but the chemical enthalpy was already released in Case 1 before reaching the reaction zone. Therefore, plasma assisted  $CH_4$  oxidation is more effective in increasing the extinction limit than  $CH_4$  reforming.

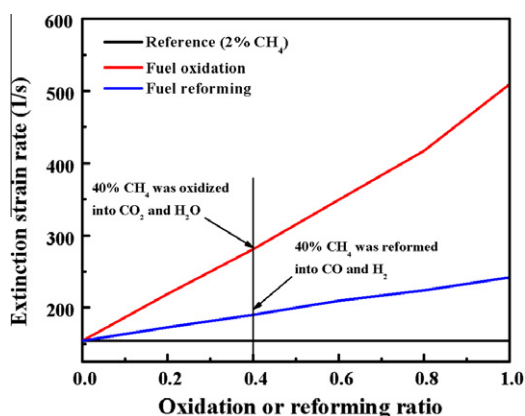


Fig. 10. Comparison of extinction strain rates of fuel oxidation and fuel reforming effect under a constant enthalpy condition for the oxidizer stream of  $Ar/He/O_2/CH_4$  (0.32/0.4/0.26/0.02 by volume),  $P = 60$  Torr.

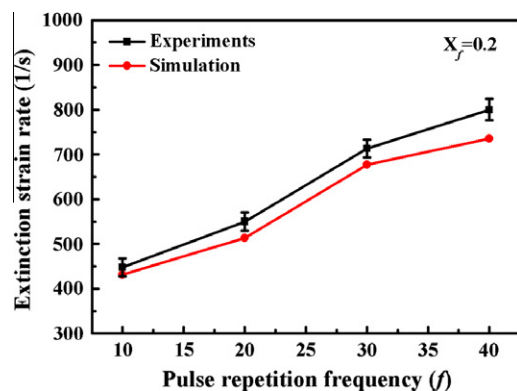


Fig. 11. Dependence of extinction strain rate on pulse repetition frequency at a fuel side  $CH_4$  mol fraction of  $X_f = 0.2$ , an oxidizer side volumetric composition of  $Ar/He/O_2/CH_4$  (0.32/0.4/0.26/0.02), and  $P = 60$  Torr.

The extinction strain rate measurements and simulations were also conducted at  $f = 20, 30$  and  $40$  kHz (with oxidizer temperature  $T_o = 743 \pm 5$  K,  $843 \pm 7$  K and  $933 \pm 8$  K, the oxidizer side flow rates were 1.66 SLPM, 1.89 SLPM and 1.92 SLPM, respectively, and with measured and calculated boundary conditions as described above) by fixing the fuel mole fraction and are shown in Fig. 11. With the further oxidation of  $CH_4$  (increased  $f$ ) in the partially premixed oxidizer stream, the extinction strain rates increased significantly. With further increase of  $f$ , the experimental and simulation results deviated, which may indicate that there are additional reaction paths to enhance the flame extinction. Nonetheless, the maximum deviation was small at approximately 10%. Moreover, the peak concentration of  $H_2$  was 1751 ppm at  $f = 20$  kHz, and the extinction limit enhancement from  $H_2$  was only 5.3%. Thus, the dominant enhancement mechanism for the extinction limit is plasma-induced prompt  $CH_4$  oxidation.

#### 4. Conclusion

The present experiments employing a nanosecond plasma discharge in a partially premixed counterflow flame showed that plasma assisted  $CH_4$  oxidation at low temperature (60 Torr) can extend the extinction limits dramatically. It was found that the main cause of the combustion enhancement was due to a kinetic-thermal effect of prompt  $CH_4$  oxidation and heat release by plasma generated species. Although plasma assisted combustion led to partial  $CH_4$  reforming to  $H_2$  and  $CO$ , numerical simulations showed that the effect of  $CH_4$  reforming on combustion enhancement was mainly through the rapid  $H_2$  chemistry but only played a minor role compared to the prompt  $CH_4$  oxidation.

The results also showed that  $O$  production by the plasma discharge was the main contributor to the prompt  $CH_4$  oxidation at low temperature and that partial premixing of fuel in the oxidizer stream led to a dramatic decrease of  $O$  concentration due to its rapid consumption by  $CH_4$  oxidation. The path flux analysis showed that  $O$  was generated primarily by the discharge via direct electron impact dissociation of  $O_2$  and  $O_2$  collisions with electronically excited species. In addition, the results also showed that the electron impact and excited species collisional dissociation of fuel also contributed approximately a mole fraction of 0.1 of the radical production.

The extinction strain rates from experimental measurements agreed well with the numerical simulations at low pressure. Comparison of measured and predicted species concentrations in plasma assisted  $CH_4$  oxidation showed that the low temperature plasma combustion kinetic model reasonably predicted the  $O$  and  $CH_2O$  formation, but the model over-predicted the concentrations of  $CO$ ,  $H_2O$ ,  $H_2$  and under-predicted the concentration of  $CO_2$ . In order to improve the predictability of the plasma flame kinetic model, the elementary reaction rates of  $O$  production by electron impact and excited species collision need to be revisited. In addition, more experimental validation data in flow reactors of  $CH_4$  oxidation at temperatures between 300 K and 700 K with radical addition are necessary.

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